Chapter 2.2 Part 1

Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Hr. \_\_\_\_

1. Section 2.2 Solutions and Suspensions
2. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ ---- composed of \_\_\_\_\_\_\_\_\_ or \_\_\_\_\_\_\_\_\_ elements or compounds that are \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ (NOT \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ ) mixed together. Examples: \_\_\_\_\_\_\_ and pepper; \_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
3. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ ----- all components are \_\_\_\_\_\_\_\_\_\_\_ distributed throughout the \_\_\_\_\_\_\_\_\_\_\_\_\_\_. \_\_\_\_\_\_\_\_\_\_\_\_ is the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ -----the substance in which the \_\_\_\_\_\_\_\_\_\_\_\_ dissolves.
4. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_-----a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of water and non \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ material. Example: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ is mostly water with \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ compounds, but it also contains \_\_\_\_\_\_\_\_ and other undissolved \_\_\_\_\_\_\_\_\_\_\_\_ that remain in \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
5. Acids, Bases, and pH

H2O H+  + OH-

----Water molecules sometimes \_\_\_\_\_\_\_\_\_\_\_ apart to form \_\_\_\_\_\_\_\_.

1. The pH Scale -----\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ system used to \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ the concentration of H+ ions in a \_\_\_\_\_\_\_\_\_\_\_\_.
2. Scales range from \_\_\_\_\_\_ to \_\_\_\_\_\_. pH 7 is \_\_\_\_\_\_\_\_\_\_\_\_\_\_, equal concentration of H+ ions and OH-  ions.
3. Acids --- any \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ that forms H+ ions in \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. Range is \_\_\_\_\_\_\_\_\_\_\_\_ with \_\_\_\_\_\_\_\_\_\_\_\_ pH = \_\_\_\_\_\_\_\_\_\_\_\_\_ acids.
4. Bases ---- any \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ that form OH-  ions in \_\_\_\_\_\_\_\_\_\_\_\_\_\_. Range \_\_\_\_\_\_\_\_\_\_\_ with higher pH = \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ bases.

Draw and label the pH scale in the space below:

**pH Scale**